Review of: "An Experimental Method to Calculate Average Metal Ions Charge by Electrolysis at Different Temperatures"

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Potential competing interests: No potential competing interests to declare.

The author claim that the calculated the average charge for cations was performance on concentrated sodium chloride (NaCl) solution. What is the concentration of this brine? what is the ionic strength? These data can explain the ion-pair formation, which result from electrostatic interaction between ions. Then this can be explaining that the charge is not an integer. Copper ions equals +1 is not usually, and this oxidation state is formed in complexes on copper. For example, thiourea reacts with cupric ions and forms a complex with Cu(I). Is not clear this state of oxidation of copper.

In the section “Experimental results are as follows”, please made tables in order to make more readable information. A well-designed table will enable the informatio to be understood.

The comparison between experiments is difficult because the electrolysis currents applied to the cylindrical anodes were different, why?

I think the conclusions need to rewrite with basis in equations from (2) to (8) and the corrosion rate mpy or mm/y